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Logged in as Quantum Dom Sciencemadness Discussion Board * Fundamentals * Organic Chemistry * Methyl Phenyl Butanone to Printable Version Subscribe Add to Favorites Author: Subject: Methyl Phenyl Butanone to Methyl Benzyl Ketone evil_lurker International Hazard Posts: 762 Registered: 12-3-2005 Location: United States of Elbonia Member Is Offline Moderator Nicodem Vorte Protest ref Prote on 14-8-2006 at 12:15 Courte Courte Registered: 12-3-2005 Location: United States of Elbonia Member Is Offline Dested on 14-8-2006 at 12:15 Moderator Prote on 14-8-2006 at 12:15 Courte Rest	\sim T	he <i>art</i> and <i>science</i> of			
Search Stack Eack to: sciencemadness.org Sciencemadness Discussion Board » Fundamentals » > Organic Chemistry Organic Chemistry w Methyl Phenyl Butanone to > Organic Chemistry + + + Methyl Benzyl Ketone > Organic Chemistry + + + Printable Version Subscribe Add to Favorites > New Topic Post Reply Author: Subject: Methyl Phenyl Butanone to Methyl Benzyl Ketone = 000000000000000000000000000000000000	Logged in as Quantum_D Logget in as Quantum_D [Logout - U2U - User Control Par				
Sciencemadness Discussion Board » Fundamentals » Organic Chemistry » Methyl Phenyl Butanone to Methyl Benzyl Ketone > Organic Chemistry Printable Version Subscribe Add to Favorites New Topic Post Reply Author: Subject: Methyl Phenyl Butanone to Methyl Benzyl Ketone evil_lurker International Hazard posted on 14-8-2006 at 12:04 QUOTE REPORT Methyl Phenyl Butanone to Methyl Benzyl Ketone Ijust managed to stumble onto this reaction in which benzaldehyde is condensed with MEK to form methyl phenyl butanone. Ijust managed to stumble onto this reaction in which benzaldehyde is condensed with MEK to form methyl phenyl butanone. Posts: 762 Registered: 12-3-2005 Location: United States of Elbonia Member Is Offline Ijust managed to stumble onto this reaction in which benzaldehyde is condensed with MEK to form methyl phenyl butanone. Moderator The ketone is then subjected to a baeyer-villiger type oxidation using a peracid, such as performic, to the corresponding ester then hydrolyzed to form methyl benzyl ketone. Moderator The ketone any references or information relating to this reaction? Moderator Vool I ILL Nicoclem posted on 14-8-2006 at 12:15	🔎 Search 1 FAQ 😫 Member List 🔂 Today's Posts 💙 Stats 🛛 🛛 Back to: scier				
Printable Version Subscribe Add to Favorites New Topic Post Reply Author: Subject: Methyl Phenyl Butanone to Methyl Benzyl Ketone posted on 14-8-2006 at 12:04 Reprovementational Hazard Methyl Phenyl Butanone to Methyl Benzyl Ketone Iternational Hazard Methyl Phenyl Butanone to Methyl Benzyl Ketone Methyl Phenyl Butanone to Methyl Benzyl Ketone Iternational Hazard Methyl Phenyl Butanone to Methyl Benzyl Ketone Iternational Hazard Methyl Phenyl Butanone to Methyl Benzyl Ketone Iternational Hazard Methyl Phenyl Butanone. This managed to stumble onto this reaction in which benzaldehyde is condensed with MEK to form methyl phenyl butanone. The ketone is then subjected to a baeyer-villiger type oxidation using a peracid, such as performic, to the corresponding ester then hydrolyzed to form methyl benzyl ketone. Dees anyone have any references or information relating to this reaction? Dees anyone have any references or information relating to this reaction? Iternational Provementation Iternational Provementation Mood: Bored out of my mind. Iternational Provementation Iternational Provementation Iternational Provementation Iternational Provementation 	Sciencemadness Discussion Board » Fundamentals » Organic Chemistry » Methyl Phenyl Butanone to Methyl Benzyl Ketone		» Organic Chemistry	***	
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*****	Nicodem Super Moderator	posted on 14-8-2006 at 12:15		QUOTE REPORT	
I assume you are looking for this:		I assume you are looking for this:			
Posts: 2042Phenylacetone by Bayer-Villiger Oxidation of 2-acetyl-1-phenylprop-1-eneRegistered:Boeeseken & Jacobs28-12-2004Recl. Trav. Chim. Pays-Bas 55, 786 (1936)Member Is OfflineFrance Content of the payor of	Posts: 2042 Registered: 28-12-2004 Member Is Offline	Phenylacetone by Bayer-Villiger Oxidation of 2-acetyl-1-phenylprop-1-ene Boeeseken & Jacobs Recl. Trav. Chim. Pays-Bas 55, 786 (1936)			
Peracetic acid treatment of 2-acetyl-propenylbenzene (3-methyl-4-phenyl-but-3-en-2-one) below 30°C resulted in the formation of phenylacetone enol acetate (2-acetoxy-1-phenylprop- 1-ene), mp 131°C and bp 103°C/3mmHg. Heating the crude reaction mixture with aqueous HCI at 80°C gave phenylacetone.		Peracetic acid treatment of 2-acetyl-propenylbenzene (3-methyl-4-phenyl-but-3-en-2-one) below 30°C resulted in the formation of phenylacetone enol acetate (2-acetoxy-1-phenylprop- 1-ene), mp 131°C and bp 103°C/3mmHg. Heating the crude reaction mixture with aqueous HCl at 80°C gave phenylacetone.			
For more information check the old Twodogs' thread from The Hive (post No 245942).		For more information check the old Ty	wodogs' thread from The Hive (post N	o 245942).	
there is a human touch of the cultist "believer" in every theorist that he must struggle against as being unworthy of the scientist. Some of the greatest men of science have publicly repudiated a theory which earlier they hotly defended. In this lies their scientific temper, not in the scientific defense of the theory Weston La Barre PROFILE FIND U2U matagaska posted on 9-8-2007 at 00:35 QUOTE	<mark>matagaska</mark> Harmless ★	there is a human touch of the cultis as being unworthy of the scientist. So repudiated a theory which earlier they the scientific defense of the theory PROFILE FIND U2U posted on 9-8-2007 at 00:35	st "believer" in every theorist that he m ome of the greatest men of science ha y hotly defended. In this lies their scien Weston La Barre	ust struggle against ve publicly ntific temper, not in QUOTE REPORT	

Posts: 2 Registered: 26-6-2007 Member Is Offline

-	Quote:
	The Baeyer-Villiger Oxidation The reaction of the above unsaturated ketone with peracetic acid was first done by Boesken reported in Rec. Trav. Chim. 55, 786 (1936). There is some discussion of this also in Patent US3980708. Also see Organic Reactions Vols 9 43 I think. By following the directions in Patent US5670661 you will get about 35% ketone based on the weight of the unsaturated ketone used. In that patent it is suggested that by recycling a higher percentage can be achieved.
	To a 1 litre flask is added 625 ml of Glacial acetic acid and 143 grms of Sodium Perborate. To this is added 100 grms of the methyl phenyl butenone with stirring and the mixture is heated to 50°C. The mixture will heat up so care has to be taken ie cooling. However if the mix gets too cool it solidifies. Stirring and heating continued for about 6 hours. Cooled poured into 1 litre H2O and extracted with toluene or DCM. The solvent is distilled leaving a yellow oil that has a pleasant smell. This is added to 500 mls of 10% NaOH solution (50/50 H20:EtOH) and stirred for 1-2 hours extracted with toluene or DCM and distilled and the fraction boiling betweer 210-220°C collected Phenyl-2-Propanone (About 35 gms)
	In the Organic Reactions review of the Baeyer-Villiger there is a reference to the oxidation of alpha Methyl Cinnamaldehyde using H2O2 catalysed by a nitrobenzene selenic acid or something like that to give the same enol ester as above but in 90% yield

I can't get hold of Glacial Acetic Acid and need some help performing the above synth with Peracetic Acid ("Peracetic Acid - Hydrogen Peroxide and peroxyacetic acid mix. 25% w/v H202")

Anyone able to explain this to me??

Cheers

PROFILE FIND U2U

posted on 9-8-2007 at 08:06

Polverone Super Administrator

I'm sure you can find information about peracetic acid oxidations without forum members providing a "recipe" for working with this particular substrate.

Posts: 2194 Registered: 19-5-2002 Location: The Sunny Pacific Northwest Member Is Offline

PGP Key and corresponding e-mail



PROFILE E-MAIL WWW FIND V2V 🛃 AIM

Sergei_Eisenstein posted on 9-8-2007 at 08:10 National Hazard

Are you certain you cannot get hold of GAA? It sounds somewhat peculiar to me.

Peracetic acid can be made by mixing hydrogen peroxide with GAA and allowing the mixture to

Posts: 287

2 of 4

QUOTE REPORT

QUOTE REPORT

Registered: 13-12-2004 Location: Waziristan Member Is Offline Mood: training	stand for a while (several days) in the dark at RT. The concentration of peracetic acid depends, inter alia, on the strenght of your hydrogen peroxide. This is how it was done in Böseken's time. Alternatively, you may want to try performic acid or oxone (in combination with e.g. GAA).		
	damnant quod non intelligunt		
	PROFILE FIND UZU		
roamingnome National Hazard	posted on 9-8-2007 at 13:53		
	Always curious about the novelity (if that's even a word) of things		
Posts: 316 Registered: 9-9-2006 Member Is Offline	I saw on http://www.orgsyn.org/orgsyn/orgsyn/prepContent.asp?prep=cv8p0323		
	that a vaguely similar creature 3-Hydroxy-3-methyl-1-phenyl-1-butanone is too unstable to be purified by distillation, and decomposes to acetophenone and acetone.		
	An exhaustive degradation study of methyl phenyl butenone might well yield interesting intermediates in higher yield then 30%say allybenzene under Ultraviolet rays		
	I know I know shut up gnome		
	PROFILE FIND U2U		
Sergei_Eisenstein National Hazard	posted on 10-8-2007 at 00:21		
Posts: 287 Registered: 13-12-2004 Location: Waziristan Member Is Offline	That's just the retro-aldol reaction. It can be expected to occur, just as dehydration might be an option when you heat the reaction mixture. I'm not convinced that compound stability problems will arise when distilling 2-acetyl-propenylbenzene. You have an a,b-unsaturated ketone conjugated with an aromatic nucleus. These usually is a stabile type of product.		
Mood: training	damnant quod non intelligunt		
	PROFILE FIND U2U		
Klute International Hazard	posted on 10-8-2007 at 07:27		
Posts: 1182 Registered: 18-10-2006	Here is a translation of the entire 1936 Boeseken Article. There are two other similar article which I know of but there isn't any vital information not present in this one. Enjoy. Attachment: TranslationBoeseken1936.rtf (11kB) This file has been downloaded 415 times		
Location: France Member Is Offline			
	PROFILE FIND U2U		
tina_craig Harmless	posted on 20-9-2007 at 11:49		

* Posts: 4 Registered: So can sodium percarbonate be substituted for sodium perborate for this procedure? 31-3-2005 Member Is Offline PROFILE FIND U2U Post Reply New Topic » Quick Reply - [Logged in as Quantum_Dom] Disable Smilies? HTML is On Smilies are On Use signature? BB Code is On Turn BBCode off? [img] Code is On Receive e-mail on reply? 0000 😄 🤮 🥸 🙄 Post Reply More smilies **Preview Post** Sciencemadness Discussion Board » Fundamentals » » Organic Chemistry **Organic Chemistry » Methyl Phenyl Butanone to Methyl** * * * **Benzyl Ketone**

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