

Figure 3. Absorption spectrum of 1-ethyl-4-methoxycarbonylpyridinyl in MTHF: (a) $5.5 \times 10^{-3} M$, at room temperature; (b, c) $6.9 \times 10^{-3} M$, at 77°K; (d) $2.9 \times 10^{-2} M$, at 77°K before irradiation; (e) after sufficient irradiation of d with a W-lamp; (f) $6.0 \times 10^{-2} M$, at 77°K before irradiation; (g) after irradiation of f.

depends on the radical concentration, the band is characteristic of the radical association and is presumably assigned to the second charge-transfer band in consideration of the appearance of two charge-transfer bands for some pyridinium salts.⁵ The first band is almost certainly near 1.9 eV. It is thus assumed that the transformation of A_S into B_T occurred through the charge-transfer excitation of A_S , followed by the intersystem crossing to the triplet configuration. This sequence may include the dynamic reorientation of the radicals in a dimer to the positions capable of triplet transitions.

Spectral change seen in Figure 1 and the concentration dependences shown in Figure 2 may be explained by the transformation of A_S into B_T . However, the change in Figure 3 is different in intensity from that of Figure 1 for the region of 500–750 nm. This difference between two radicals will be caused by the fact that the absorptions in the 500–900-nm region originate in several kinds of species, such as radical monomer and singlet and triplet radical dimers. The monomer has an absorption due to the lowest excitation, $\Psi_g \rightarrow \Psi_1$,² in this region, which corresponds to the broad and weak absorption of the radical at room temperature.⁶ The dimers would show the similar absorption due to the local excitation in a dimer. The radical association clearly leads to the appearance of other absorptions which are dependent on the radical concentration, as seen in Figure 2. It is thus difficult to interpret in detail the absorption intensity and the assignment for the 500–900-nm region. This region also bears relation to the π -mer formation⁷ and the appearance of two charge-transfer bands demonstrated for some pyridinium salts.⁵

Further study on the transformation in the glassy state is in progress for the associations of analogous radicals.

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- A Dewar vessel with quartz windows was constructed for the measurement. The cell was immersed in liquid nitrogen.
- When the radical was purified carefully by distillation, cooling of the radical solution did not bring about the significant change of the spectrum for the region shorter than 400 nm (Figure 1a and 1b), and the absorption intensity in 650-nm region at 77°K is smaller than that reported previously.² Addition of a metal halide to the solution sometimes results in an increase of the intensity at around 610 nm with the spectral change in the shorter region. For example, $\lambda_{max}(\epsilon)$ for **1a** in MTHF with lithium iodide at 77°K is 618 nm (4800) at the concentration of $4.7 \times 10^{-3} M$. This is explained by the complex formation of **1a** with lithium iodide.¹
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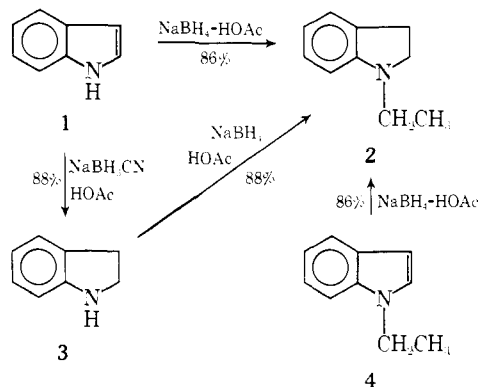
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Reactions of Sodium Borohydride in Acidic Media. I. Reduction of Indoles and Alkylation of Aromatic Amines with Carboxylic Acids¹

Sir:

We wish to report that sodium borohydride (NaBH_4) in neat carboxylic acids sequentially reduces the indole double bond and alkylates the nitrogen atom to give *N*-alkylindolines, e.g., **1** \rightarrow **2**, and that this combination of reagents conveniently alkylates primary and secondary aromatic amines, e.g., **3** \rightarrow **2** (Scheme I).

Scheme I. Transformations in Acetic Acid



Although the reduction of indoles to indolines has received considerable attention,² there is no general, efficient procedure for this transformation. Encouraged by the tendency of the indole ring to protonate at the 3-position^{3,4} and by the observations that enamines can be reduced by NaBH_4 in acetic acid-tetrahydrofuran (HOAc-THF)⁵ and sodium cyanoborohydride (NaBH_3CN),⁶ we have examined the behavior of indoles with NaBH_4 in neat carboxylic acids.

Quite unexpectedly, the reaction of indole (**1**) with NaBH_4 in glacial HOAc gives *N*-ethylindoline (**2**) in 86% yield. Likewise, the reactions of indoline (**3**) and *N*-ethylindole (**4**) with NaBH_4 -HOAc give **2** in high yield.

This unprecedented reduction-alkylation of indoles and reduction of *N*-alkylindoles⁷ with NaBH_4 in liquid carboxylic acids appears to be a general transformation (Table I). However, the stronger acid, formic, produces indole dimers and other products in addition to *N*-methylindoline. Interestingly, the reaction of **1** with NaBH_4 in trifluoroacetic acid ($\text{CF}_3\text{CO}_2\text{H}$) gives indoline (**3**), the product of reduction without alkylation, in low yield. The yield of **3** can be increased to 88% when **1** is treated with NaBH_3CN -HOAc (Table I). This latter reaction permits a very convenient synthesis of **3**.

We believe that the reaction of **1** \rightarrow **2** involves 3-protonation of indole,⁴ followed by reduction of the resulting indolenium ion to give **3**,⁸ which is subsequently alkylated (*vide infra*). The reduction of the indoloquinolizidine alkaloid **5**

Table I. Reaction of Indoles with $\text{NaBH}_4\text{-RCO}_2\text{H}^a$

Substrate	Carboxylic acid	Product ^b	Yield, % ^c
1	HOAc	<i>N</i> -Ethylindoline	86
	HCO ₂ H	<i>N</i> -Methylindoline	53
	CH ₃ CH ₂ CO ₂ H	<i>N</i> - <i>n</i> -Propylindoline	69
	(CH ₃) ₂ CHCO ₂ H	<i>N</i> -Isobutylindoline	49
	CF ₃ CO ₂ H	Indoline ^d	36
	HOAc-NaBH ₃ CN	Indoline	88
	HOAc	1-Ethyl-2-methylindoline	84
	HOAc	1-Ethyl-3-methylindoline	45
2-Methylindole	HOAc	1-Ethyl-2,3-dimethylindoline	60
3-Methylindole	HOAc	<i>N</i> -Ethylhexahydrocarbazole	77
2,3-Dimethylindole	HOAc	1-Ethyl-7-methylindoline	90
Tetrahydrocarbazole	HOAc	<i>N</i> -Methylindoline	86
7-Methylindole	HOAc	<i>N</i> -Ethylindoline	86
<i>N</i> -Methylindole	HOAc	1,2-Dimethylindoline	84
<i>N</i> -Ethylindole	HOAc		
1,2-Dimethylindole	HOAc		

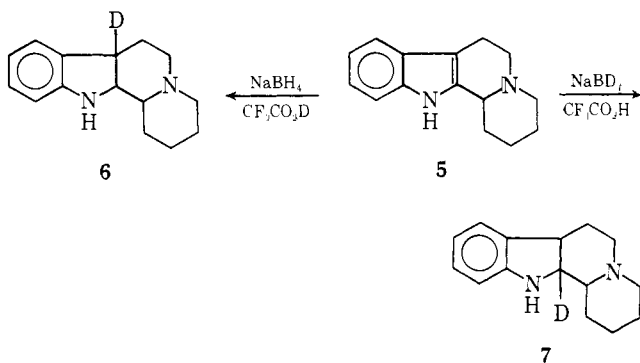
^a Conditions are typically 0.005–0.02 mol of indole dissolved in 30–150 ml of dry carboxylic acid at 15–20° to which is slowly added 0.02–0.2 mol of NaBH_4 pellets (Alfa Inorganics, Inc., Beverly, Mass.). A brief heating period at 50–60° is sometimes necessary to complete the reaction. Reactions can be monitored by tlc or uv. ^b Identified by comparison with authentic material or by conversion to known derivatives (picrate, methiodide). All products exhibited satisfactory ir, nmr, and uv spectra. ^c Yields are for pure distilled material. In most cases starting indole was the only other material present in the reaction mixture. The reactions have not been optimized. ^d A small amount of *N*-tri-fluoroethylindoline is also formed.

Table II. Reaction of Aromatic Amines with $\text{NaBH}_4\text{-RCO}_2\text{H}^a$

Substrate	Carboxylic acid	Product ^b	Yield, % ^c
Aniline	HOAc	<i>N</i> -Ethylaniline	88
	HOAc (50–60°)	<i>N,N</i> -Diethylaniline	74
	HOAc-Me ₂ CO	<i>N</i> -Isopropylaniline	68
	HOAc-Me ₂ CO (50–60°)	<i>N</i> -Ethyl- <i>N</i> -isopropylaniline	79
	HOAc-PhCHO (50–60°) ^d	<i>N</i> -Benzyl- <i>N</i> -ethylaniline	80
	(CH ₃) ₂ CCO ₂ H ^e	<i>N</i> -Neopentylaniline	80
<i>N</i> -Methylaniline	HOAc	<i>N</i> -Ethyl- <i>N</i> -methylaniline	72
	HOAc-Me ₂ CO	<i>N</i> -Isopropyl- <i>N</i> -methylaniline	78
	HCO ₂ H	<i>N,N</i> -Dimethylaniline	77
	HOAc-(HCHO) _n -THF	<i>N,N</i> -Dimethylaniline	59
	CH ₃ CH ₂ CO ₂ H	<i>N</i> -Methyl- <i>N</i> -propylaniline	83
	CH ₃ CH ₂ CO ₂ H	<i>N</i> -Ethyl- <i>N</i> -propylaniline	70
<i>N</i> -Ethylaniline	HOAc	<i>N</i> -Ethyl- <i>N</i> -isopropylaniline	69
<i>N</i> -Isopropylaniline	HOAc	<i>N</i> -Ethylindoline	88
Indoline	HOAc	<i>N</i> -Trifluoroethylindoline ^f	7 ^g
Diphenylamine	HOAc	<i>N</i> -Ethyldiphenylamine	80
Carbazole	HOAc	<i>N</i> -Ethylcarbazole	92
5 <i>H</i> -Dibenz[<i>b,f</i>]azepine	HOAc	9-Ethyl-5 <i>H</i> -dibenz[<i>b,f</i>]azepine	72

^{a–c} See corresponding footnotes in Table I. ^d *N*-Benzylaniline could be isolated from the 20° reaction. ^e This reaction was also run with diglyme as a cosolvent. ^f Indoline was recovered in 51% distilled yield. ^g From the pot residue there was isolated in 34% yield an indoline dimer containing three trifluoroethyl groups.

with $\text{NaBH}_4\text{-CF}_3\text{CO}_2\text{H}$ proceeds without alkylation (72%) and the deuteration experiments, **5** → **6** and **5** → **7**, are in accord with our suggested mechanism.



The combination $\text{NaBH}_4\text{-RCO}_2\text{H}$ also provides for the facile alkylation of a variety of primary and secondary aromatic amines⁹ (e.g., **3** → **2**), and the reaction can be controlled to give mono- or dialkylation of primary amines. Thus, aniline with $\text{NaBH}_4\text{-HOAc}$ at 20° gives *N*-ethylaniline, and further reaction at 60° gives *N,N*-diethylaniline (Table II).

This alkylation method is extended by our observation that aldehydes and especially ketones are reduced to alcohols relatively slowly by $\text{NaBH}_4\text{-RCO}_2\text{H}$,¹⁰ so that unsymmetrical tertiary amines can be prepared from primary amines in one flask. Thus, aniline with $\text{NaBH}_4\text{-HOAc}$ -acetone gives either *N*-isopropylaniline or *N*-ethyl-*N*-isopropylaniline, depending on the temperature (Table II). This versatility is not available with previous methods for reductive amination^{6,11} of aldehydes and ketones (H_2 ,^{11a,b} HCO_2H ,^{11c,d} NaBH_3CN ,^{6,11e} NaBH_4 ,^{11f,g} $\text{Fe}(\text{CO})_5$,^{11h}).

We view the amine alkylation as a stepwise process: (1) reduction of carboxylic acid to aldehyde¹² (or aldehyde equivalent), perhaps *via* one or more acyloxyborohydride species¹³ and intra- or intermolecular hydride reduction of the carbonyl group; (2) reaction of the aldehyde with amine to form an iminium ion; and (3) hydride reduction¹⁴ of the iminium ion to product amine.

Although amides are side products in some cases, they are very clearly not obligatory intermediates in the alkylation reaction. Thus, *N*-acetylindoline and *N*-acetylindole are recovered in 67 and 82% yield, respectively, after treatment with $\text{NaBH}_4\text{-HOAc}$, under conditions which convert indoline completely into *N*-ethylindoline. Likewise, it seems unlikely that diborane⁷ is the reducing agent in the reaction, since externally generated gaseous diborane bubbled

into amine-HOAc gives clean acylation and *not* alkylation of the amine.¹⁵

Acknowledgment. We are grateful to the National Institutes of Health (CA-14237), Merck Sharp and Dohme, and Eli Lilly for their generous financial support of this work and our research program and to Mr. Randall B. Nelson and Dr. Jackson P. English for stimulating discussions.

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- After short reaction periods the *N*-substituted indoline can usually be detected (tlc) in the reaction mixture.
- We have also been able to alkylate several aliphatic amines with NaBH₄/RCO₂H (e.g., cyclohexylamine, benzylamine, dibenzylamine, piperidine, pyrrolidine) as part of a study which will be reported separately.
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Received August 12, 1974

Stereochemistry of the Oxythallation of 1,2-Cyclononadiene

Sir:

Although a great many studies on the mechanism of the oxymercuration reaction have been disclosed,¹ there have been comparatively few studies on the analogous oxythallation reaction. The oxythallation reaction with alkenes has been shown to have many similarities to the oxymercuration reaction.² It has been suggested that rearrangement of a π complex to a σ -bonded oxythallation adduct is rate limiting.^{2a} With Tl(III) acetate in aqueous acetic acid, the reaction was shown to be first order in alkene and thallic ion.^{2a}

A correlation of the rate of oxidation of alkenes by Tl³⁺ with Brown's σ^+ values ($\rho^+ = 2.2$)^{2b} and Taft's σ^* values ($\rho^* = -3.2$)^{2c} suggests that there is a high degree of positive charge delocalization in the transition state. In the latter study no kinetic evidence for a thallium ion intermediate was found.

The most notable difference between the two oxythallation reactions is the lack of solvolytic stability of the carbon-thallium bond resulting in dethallation of the organothallium intermediates. In fact, only a few oxythallation adducts of olefins have been isolated. The oxythallation products of styrene³ and isobutylene⁴ suggested that the reaction proceeds in the Markownikov sense. Because of the ease of heterolysis of the C-Tl bond, conclusive evidence for the stereochemistry of this electrophilic addition reaction is still lacking. Norbornene and norbornadiene have been shown to afford *exo cis* adducts on acetoxythallation.⁵ An anti mode of addition to 4-*tert*-butylcyclohexene has been inferred on the basis of indirect evidence.⁶

We chose 1,2-cyclononadiene (**1**)⁷ as a model compound to elucidate the stereochemistry of oxythallation since the vinyl thallium adducts⁸ are readily isolated and characterized. Solvolysis of the C-Tl bond in **4** would afford a relatively unstable vinyl cation. This study also affords a direct comparison with the oxymercuration of **1** which has been thoroughly investigated.^{9,10} The mechanism of the oxythallation of **1** is of particular interest since the oxymercuration of this cyclic allene proceeds by an anti mechanism¹⁰ while the acetoxyplumbation occurs in a *syn* fashion.¹⁰ We now report that oxythallation of optically active (**1**)⁷ affords optically active products by an anti addition.

Treatment of **1** with an equivalent of thallic acetate in glacial acetic acid afforded the oxythallation adduct **4a** (84%).¹¹ Reduction of **4a** with basic NaBH₄ afforded *cis*-3-acetoxycyclononene (**5a**) (61%) that was identical in every respect to an authentic sample prepared by the acetoxymercuration-demercuration¹⁰ of **1**. Acetoxythallation of optically active **1**, [α]^{25D} -15.6° afforded **4a** that had [α]^{25D} *ca.* -0.5° which on reduction with NaBH₄ afforded **5a**, [α]^{25D} +0.6°. The reaction of **1** with Tl(OAc)₃ in methanol also afforded a stable methoxythallation adduct **4b**¹² (76%). Demetalation with NaBH₄ gave *cis*-3-methoxycyclononene (**5b**)¹⁰ (72%). The position of the diacetatothallium moiety was further established by treatment of **4b** with Br₂ in CCl₄ affording *cis*-2-bromo-3-methoxycyclononene (**6**)¹⁰ (82%). Methoxy- and ethoxythallation-dethallation of optically active **1** also afforded optically active allylic ethers (*S*)-(+)-**5** (Table I).

The isolation of an optically active product from these reactions establishes that the planar allylic cation **3** cannot be the sole precursor to **4**. Thus, bridging due to π -complex formation is sufficient to prevent complete carbon-carbon bond rotation affording **3**. The oxythallation of (*S*)-(-)-1,2-cyclononadiene^{10,13} to afford (*S*)-(+)-3-acetoxy- and alkoxycyclononene¹⁴ must proceed by an anti mode of addition. Our data also show that the relative stereospecificity of the oxythallation reaction is comparable to that of the anti oxymercuration reaction but is considerably less than that observed for the *syn*-acetoxyplumbation of **1**.¹⁵ The optical purity of **5b** and **5c** was further reduced when Tl(NO₃)₃ was used (Table I).¹⁶

We have also established that the oxythallation adducts **4b** and **4c** (X = OAc or NO₃) are not formed reversibly from either **2** or **3**. Thus, treatment of **4b** in EtOH solvent or **4c** in CH₃OH solvent in the presence of an equivalent of HNO₃ did not result in alkoxy exchange. Similarly, attempts to exchange alkenes by treatment of **4a** (X = NO₃) with 1-octene did not effect an alkene exchange. In contrast, reaction of the methoxymercureal of 1-octene (X =